


## Research Article

# Functionalized Graphene Derivatives: Antibacterial Properties and Cytotoxicity

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In this work, the authors prepared and characterized two different graphene oxides: one chemically synthesized (GO sample) and the other one electrochemically synthesized ( $\text{GO}_{(\text{LiCl})}$ ). Both samples were fully characterized with atomic force microscopy (AFM), Raman and Fourier transform infrared (FTIR) spectroscopies, X-ray photo electron spectroscopy (XPS), thermal analysis (TG/DTA), and Z-potential. The antibacterial properties of both graphene oxides were studied using Gram-negative *Escherichia coli* ATCC 25922 and Gram-positive *Staphylococcus aureus* ATCC 25923 by spectrophotometer and viable cell count as indirect and direct methods, respectively. Results demonstrated that the  $\text{GO}_{(\text{LiCl})}$  exhibited a significant antibacterial activity compared to GO that showed a bacteriostatic effect on both pathogens. Electron microscopy analysis confirmed the antibacterial effects of both graphene oxides toward the pathogens, especially working at 80  $\mu\text{g}/\text{mL}$ , for 24 h. Additional studies were also performed and both GO samples were not cytotoxic at 2  $\mu\text{g}/\text{mL}$  toward neuroblastoma cells. Moreover, 2  $\mu\text{g}$  of GO was suitable to carry the minimum effective dose (5.74 ng/mL) of kinase inhibitor S29 (1-(2-chloro-2-(4-chlorophenyl)ethyl)-N-(4-fluorobenzyl)-1H-pyrazolo[3,4-d]pyrimidin-4-amine), providing negligible side effects related to the S29 treatment (this latter being specifically active on the neuroblastoma cell lines (SK-N-BE(2))).

## 1. Introduction

One of the most important research field applications for medicine focuses on new therapeutic agents' delivery tools, with the first aim to improve therapy efficiency. Recent advances in the nanomaterial field provide useful biocompatible nanostructured materials and nanocomposites, suitable for drug delivery systems. Among all the innovative materials, graphene derivatives [1] have attracted a great deal of scientific interest in biomedicine and pharmacology

sciences. Graphene, a single layer of  $\text{sp}^2$ -hybridized carbon atoms arranged in a honeycomb two-dimensional (2-D) crystal lattice, exhibits excellent chemical-physical, biochemical [2], mechanical, and engineering properties, extremely useful for several applications in sensors [3] and optoelectronic devices [4], environmental monitoring tools [5], food quality control [6], and medical devices [7]. With particular regard to the fields of application that include food quality control and medicine field applications, the antibacterial features [8] and the biocompatibility properties toward the real

matrixes (as food or normal human cell lines, organic tissue, and blood) represent a key point that nanomaterials must possess. There are several cases of study, described in literature by interesting review articles concerning the improved antibacterial activity [9] and the increased biocompatibility [10] properties, exhibited by the graphene derivatives. These last properties mainly depend on the synthesis routes applied for graphene production [11]. According to this, Valentini et al. [12] recently report about a metal-free green graphene oxide ( $\text{GO}_{(\text{LiCl})}$ ), synthesized by the electrochemical exfoliation of micrometric graphite electrodes (2015 patent no. 102015000023739).  $\text{GO}_{(\text{LiCl})}$  was fully characterized, resulting in antibacterial material, and finally applied for in vitro cellular experiments, to establish its biocompatibility [12]. The novelty of this work consists of a new green chemistry synthetic route (under patent) for the fabrication of GO samples. The latter showed a bacteriostatic activity (if compared with the bactericidal action, exhibited by  $\text{GO}_{(\text{LiCl})}$ ), in the presence of both pathogens, as the Gram-negative *E. coli* ATCC 25922 and the Gram-positive *S. aureus* ATCC 25923. A cytotoxicity assay was performed, in order to identify in GO a new promising nanocarrier for compound S29: 1-(2-chloro-2-(4-chlorophenyl)ethyl)-N-(4-fluorobenzyl)-1H-pyrazolo[3,4-d]pyrimidin-4-amine, therapeutic agent. The scientific interest on S29 as a therapeutic agent consists of its ability to reduce tumor mass in neuroblastoma xenograft mice models. For this purpose, in this study the SK-N-BE(2) neuroblastoma cell lines were used for the cytotoxicity investigation, concerning the S29 therapeutic anticancer agent. S29 shows an unfavorable pharmacokinetic profile, causing a limited use in the treatment of cancer. Until now, in literature, liposomes [13] and human serum albumin (HSA) [14] have been applied to improve S29 delivery, biodistribution, and bioavailability but the concentration of liposomes and HAS nanoparticles had to be very high and this leads to significant high cytotoxicity in vitro and in vivo [13, 14] experiments. Instead, in this work, it has been shown that graphene derivatives were not cytotoxic and they could be used at extremely low levels of concentration, if compared with those applied in the presence of liposomes and HAS nanoparticles. In particular, according to the results exhibited by the GO samples, this latter seems to be more eligible for S29 delivery.

## 2. Materials and Methods

**2.1. Chemicals.** The electrochemically grown graphene oxide sample (labelled as  $\text{GO}_{(\text{LiCl})}$ ) was obtained by a patented synthetic approach, as widely reported in a 2015 patent (no. 102015000023739) and also in reference [12]. For the chemical synthesis of the graphene oxide (labelled as GO), a new wet mechanic-chemical exfoliation synthesis was developed by the same authors of this work, and now it is under patent. The other chemicals used in the analysis were purchased from Sigma-Aldrich (USA). All reagents were of analytical grade and used as received without further purification.

The procedure of preparing GO and  $\text{GO}_{(\text{LiCl})}$  nanosheet dispersions was similar for the cytotoxicity and antibacterial

tests, except that GO suspension was carried out by using phosphate-buffered saline (PBS) for the cytotoxicity analysis and by using a saline solution 0.9% (10 mL) in the case of the antibacterial tests. The final graphene oxide dispersions were transparent, and they were very well dispersed by sonication with a polytronic probe, at room temperature.

**2.2. Material Characterization.** The AFM measurements were performed in air using a Veeco Multiprobe IIIa instrument (Veeco, Santa Barbara, CA, USA). The GO was deposited using an aqueous solution on the mica substrates. The experiments were carried out at room temperature (20°C) in tapping mode using Si tips with a spring constant of about 40 N/m and a typical curvature radius on the tip of 7 nm. The XPS measurements were carried out using an Omicron DAR 400 Al/Mg  $K\alpha$  nonmonochromatic X-ray source and a VG-CLAM2 electron spectrometer. GO (1 mg/mL) was very well dispersed in ethanol using an ultrasonic probe and then deposited on the silicon wafer.  $\zeta$ -potential was determined by Dynamic Light Scattering (DLS) (Zetasizer Nano ZS90, Malvern Instruments Ltd., Malvern, UK). Graphene oxide samples were diluted in water, providing an optical density (OD) of 0.1 at 420 nm, and then measured at 24°C with a scattering angle of 90°.  $\zeta$ -potential measurements were carried out using folded capillary cells (Malvern). Thermogravimetric analysis was carried out by a Q600 thermogravimetric analyzer (TA Instruments-Waters, USA). Samples (~20 mg) were put in a platinum crucible and heated from 30°C to 700°C, with a rate of 10°C/min, under nitrogen purge gas.

**2.3. Cytotoxicity Tests: Cell Culture.** The long-term neuroblastoma cell lines, SH-SY5Y and SK-N-BE(2) (ACC no. 209 and ACC no. 632, respectively), obtained from DSMZ, Braunschweig, Germany, in October 2012, were maintained in RPMI 1640 (Gibco, Grand Island, NY, USA) with or without phenol red, supplemented with heat-inactivated 10% FCS containing 2 mM L-glutamine. Preliminary time and concentration course experiments with GO 0.57, 5.74, and 57.4  $\mu\text{g/mL}$  were performed in our laboratory in order to establish the conditions at which GO did not cause cell necrosis. Cell vitality was monitored before each experiment by a trypan blue (Sigma-Aldrich) [15] exclusion assay, in which dead cells are stained and those with intact membranes are not. Briefly, cell viability was determined using the trypan blue exclusion test. SK-N-BE(2) cells ( $5 \times 10^4/\text{mL}$ ) were suspended in a 10  $\mu\text{L}$  suspension containing trypan blue (0.4% w/v; Sigma-Aldrich) in phosphate-buffered saline (PBS) and incubated at room temperature for 5 min, followed by examination under a light microscope to determine the percentage of cells with clear (viable cells) and blue (nonviable cells) cytoplasm.

**2.4. Antibacterial Activity.** GO or  $\text{GO}_{(\text{LiCl})}$  (2 mg) were dispersed in saline solution 0.9% (10 mL) by sonication with a polytronic probe. The suspensions were sterilized by autoclaving, at 121°C for 15 min, before being inoculated with bacteria to reach a final concentration ranging from 20 to 80  $\mu\text{g/mL}$ . *Staphylococcus aureus* subspecies aureus derived from ATCC 25923 and *Escherichia coli* derived from

TABLE 1: Chemical-physical parameters and microanalysis of the graphene oxide samples.

Sample	Shape	Thickness (nm)	Area ( $\mu\text{m}^2$ )	Microanalysis % (w/w)
Chemically synthesized GO	Rectangular nanosheets	1.5 $\pm$ 0.4 No. layers: 2-3	$6 \times 10^{-2}$ - 0.1	Si, S, Ca, Cr, Fe, Co n.d.
Electrochemically synthesized GO	Rectangular nanosheets	1.2 $\pm$ 0.3 No. layers: 2	0.1-3.0	Si, S, Ca, Cr, Fe, Co n.d.

n.d.: not detectable.

ATCC 25922 were purchased from Microbiologics (USA). Tryptic Soy Broth (TSB) was purchased from VWR Prolabo Chemicals (Italy) and prepared as reported in data sheet. Nutrient Agar with NaCl was purchased from BioLife (Italy) and prepared as reported in product information. Saline solution 0.9% (NaCl: VWR Prolabo Chemicals, Italy) was prepared dissolving 9 g of NaCl in 1000 mL of distilled water, sterilized by autoclaving at 121°C for 15 min. *E. coli* (ATCC 25922) and *S. aureus* (ATCC 25923) were used to perform viability tests. Bacteria were grown in TSB medium at 37°C overnight, and cells were collected via centrifugation (4000 rpm for 20 min). The cells were washed three times with sterile saline solution to remove residual macromolecules and other growth medium constituents. The pellets were then suspended in saline solution at 0.9% [16, 17]. Bacterial cell suspensions were diluted in NaCl 0.9% to obtain cell samples having an optical density at 600 nm ( $\text{OD}_{600}$ ) at  $T_0$  corresponding to  $10^7$ - $10^8$  CFU/mL ( $\text{OD}_{600} = 0.4$  for *S. aureus* and  $\text{OD}_{600} = 0.3$  for *E. coli*). Cells were incubated with different concentrations (20, 40, and 80  $\mu\text{g}/\text{mL}$ ) of GO or  $\text{GO}_{(\text{LiCl})}$  at 37°C for 4 hours ( $T_1$ ) and 24 hours ( $T_2$ ). The untreated samples with *E. coli* or *S. aureus* were used as the control. Growth of treated and untreated samples were followed by measuring OD at 600 nm with spectrophotometer NANOCOLOR UV/VIS II (MACHEREY-NAGEL, Germany), at 1-hour intervals for 6 hours and then after 24 hours of treatment. The viability was evaluated by the colony-counting method. Briefly, bacterial amount at the starting time ( $T_0$ ), after 4 hours of treatment ( $T_1$ ), and after 24 hours ( $T_2$ ) was determined by performing a series of 10-fold cell dilutions (100  $\mu\text{L}$  each) that were spread onto Nutrient Agar plates and allowed to grow overnight at 37°C. Colonies were counted and compared with those on control plates to calculate the change in cell growth inhibition. The antibacterial activity was expressed as a function of cell viability loss; the loss of cell viability (%) was determined by using the number of colonies found in the experimental mixture incubated with GO and the number of colonies found in the “inoculum only,” incubated without GO. All the experiments were done in duplicate, and the average values were reported. For SEM (FE-SEM Quanta Inspect F, FEI, Thermo Fisher Scientific) analysis, a steady state culture of *E. coli* (ATCC 29522) and *S. aureus* (ATCC 29523), grown in LB at 37°C o.n., was washed in physiological saline solution and incubated with GO or  $\text{GO}_{(\text{LiCl})}$  80  $\mu\text{g}/\text{mL}$ , for 8 and 24 hours at 37°C. Successively, samples were fixed with 2.5% glutaraldehyde in 0.1 M sodium cacodylate buffer and were left to adhere on poly-lysine-treated glass coverslips for 4 hours at room temperature (RT) and

TABLE 2: Binding Energies (BE) and deconvoluted peaks (%) for C1s of GO.

Peak BE (eV)	C1s At. %	Functional groups
284.4	59.0	C-C
285.4	16.0	C-OH
286.6	14.0	C-O
287.7	7.0	C=O
289.0	4.0	C(O)O
290.7	2.2	$\pi$ - $\pi^*$

$\pi$ - $\pi^*$ : means the  $\pi$ -bonding orbital and the  $\pi^*$ -antibonding orbital.

TABLE 3: Binding Energies and deconvoluted peaks (%) for C1s of  $\text{GO}_{(\text{LiCl})}$ .

Peak BE (eV)	C1s At. %	Functional groups
283.6	23.50	C-C
284.6	36.50	C-OH
285.5	21.70	C-O
286.5	18.30	C=O

o.n. at 4°C. Then, adhered bacteria were washed and post-fixed with 1%  $\text{OsO}_4$  in 0.1 M sodium cacodylate buffer for 1 h at RT and dehydrated through a graded series of ethanol solutions (from 30% to 100%), according to reference [18]. For TEM (EM 208S transmission electron microscope, FEI, Thermo Fisher, equipped with the acquisition system Olympus MegaView SIS camera, at 100 kV) analysis, the same bacterial cultures were fixed in 2.5% glutaraldehyde, 2% paraformaldehyde, and 2 mM  $\text{CaCl}_2$  in 0.1 M sodium cacodylate buffer, pH 7.4, overnight at 4°C, and were processed according to Perry and Gilbert [19]. Samples were washed in cacodylate buffer and postfixed with 1%  $\text{OsO}_4$  in 0.1 M sodium cacodylate buffer for 1 hour at RT, treated with 1% tannic acid for 30 min and rinsed in 1% sodium sulphate, for 10 min. Postfixed specimens were washed, dehydrated through a graded series of ethanol solutions (30–100% ethanol), and embedded in Agar 100 (Agar Scientific Ltd., UK).

### 3. Results and Discussion

**3.1. Characterization of GO Samples.** These two graphene oxide samples were synthesized by a chemical and electrochemical synthesis pathway, respectively, and fully characterized by using AFM, EDAX, and XPS techniques. The

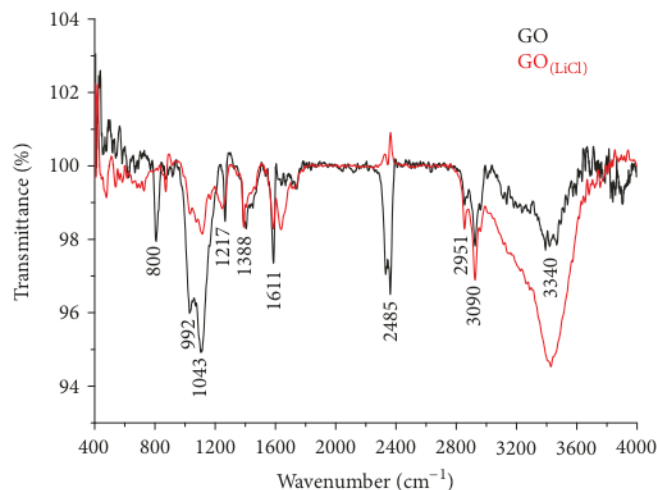


FIGURE 1: FTIR spectra profiles for both GO derivatives.

TABLE 4: Physical properties of the graphene oxide samples.

Sample	Weight loss % (TGA) <sup>a</sup>	Acidic sites (nmol/mg) <sup>b</sup>	Extent of defects ( $I_D/I_G$ ) <sup>c</sup> , Raman
GO	$0.88 \pm 0.23$	$9.45 \pm 0.20$	0.42
GO <sub>(LiCl)</sub>	$0.42 \pm 0.40$	$4.02 \pm 0.23$	0.10

<sup>a</sup>Mean  $\pm$  SD. <sup>b</sup>Acidic sites evaluated as described in [29, 30] (Mean  $\pm$  SD), by volumetric titration. <sup>c</sup>Evaluated by Raman spectroscopy.

main characterization results have been highlighted in Tables 1, 2, and 3. In Table 1, it is possible to recognize that the two kinds of graphene oxide samples were very similar in thickness and number of layers; on the other hand, chemically synthesized graphene (GO) had a lower surface area value, if compared with that exhibited by the electrochemically synthesized graphene sample (GO<sub>(LiCl)</sub>). Instead, as regards the content of metals, the microanalysis (EDAX) unequivocally showed that the metal content resulted below the detection limit of the microanalysis technique, indicating that both these GO samples were metal-free graphene-based nanomaterials.

XPS analysis was carried out to study the chemical shift of GO samples. The C1s XPS core level spectrum analysis of GOs with the corresponding quantitative parameters is shown in Table 2 (for GO) and Table 3 (for GO<sub>(LiCl)</sub>). From Table 2, it is possible to observe that the chemically synthesized graphene oxide sample results are more oxidized, with a greater presence of carbonyl and/or carboxylic functional groups, if compared with those exhibited by the graphene synthesized by the electrochemical route strategy (see also Table 3). The electrochemically synthesized graphene sample results in a more reduced C1s chemical shift (i.e., less oxidized if compared to the chemical shift of graphene obtained by chemical synthesis approach).

The functionalization degree and the chemical shift of graphene oxide samples certainly represent very useful

parameters, strictly related to the antibacterial activity and biocompatibility properties. According to the literature [20], the oxygenated functional groups are firstly responsible for the antibacterial behavior of GO nanosheets. The mechanism underlying this feature involved the generation of reactive oxygen species (ROS), able to provoke the oxidative stress through an electronic transfer mechanism. Functionalities strictly influence the electronic and/or electron-transfer processes, especially since graphene oxide shows a more insulating profile, when compared with pristine (metallic behavior) and reduced graphene sheets (semiconductor-semimetal transitions with reduction). The electronic/electron-transfer interaction mechanisms mainly occur between the positively charged Gram-negative *E. coli* and the negatively charged GO and GO<sub>(LiCl)</sub>, preferentially. In the case of the Gram-positive *S. aureus* (having negatively charged bacterial walls), the semiconductive properties of GO<sub>(LiCl)</sub> could provoke a greater electrostatic repulsion (when compared to the insulating GO), because it delivers a greater electron density, toward the negatively charged *S. aureus* pathogen walls (and the electron-transfer mechanism is discussed in detail in Results and Discussion, Section 3 of this paper).

Several authors affirmed that the antibacterial features of nanomaterials are also the result of a complex interaction between two main additional factors: size and shape. Liu et al. [21] demonstrated that the antibacterial activity of GO sheets toward *E. coli* cells is lateral size dependent. Larger GO sheets show stronger antibacterial activity than smaller ones, and they have different time- and concentration-dependent antibacterial features. The same authors [21] described that the lateral size-dependent antibacterial properties of GO sheets are caused by neither their aggregation states nor oxidation ability. Large GO sheets more easily cover cells, and cells cannot proliferate once fully covered, resulting in the cell viability loss, experimentally observed in the following colony-counting test. In contrast, small GO sheets adhere to the bacterial surfaces, which cannot effectively isolate cells from the environment.

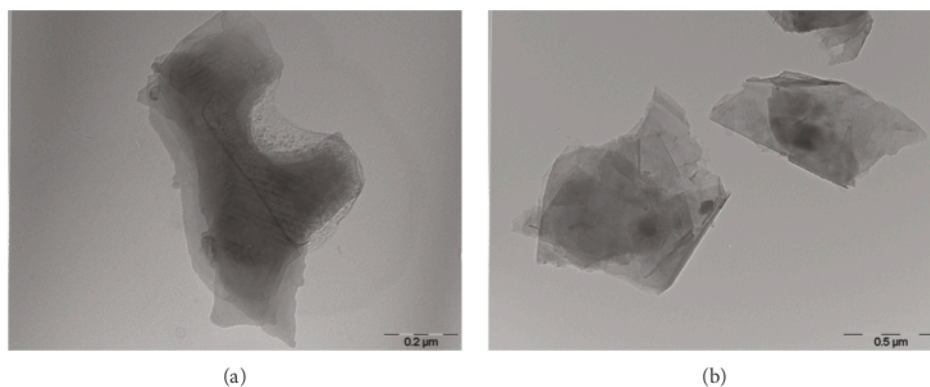


FIGURE 2: TEM micrographs of (a) GO and (b)  $\text{GO}_{(\text{LiCl})}$ .

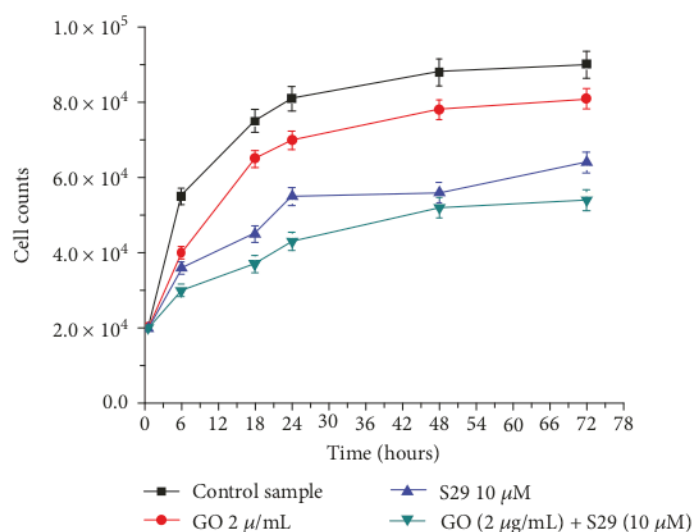


FIGURE 3: Growth curves of SK-N-BE(2), treated with GO (2.0 μg/mL) and S29 10 μM, 72 h. RSD < 3%.

As far as the form is concerned, it is stated in the literature that the geometries/shapes of both nanomaterials and bacteria are fundamental to guarantee the mutual interaction of the two systems. In particular, Akhavan et al. [22] affirmed that the graphene nanoribbons showed much higher cytotoxicity than the graphene sheets. The ROS (Radical Oxygenated Species) generation and direct contact interaction of the extremely sharp edges of graphene with the membrane of the cells were found as the main mechanisms involved in the cytotoxicity of both shapes of graphene: nanoribbons and nanosheets [22].

Moreover, the shape of the bacteria is also relevant. Hui et al. [23] report about diverse Graphene Quantum Dots (GQDs) with basal planes similar to those of graphene oxide sheets lacking an antibacterial property, which is prepared by rupturing the C60 cage which effectively kills *Staphylococcus aureus*, including its antibiotic-tolerant persisters, but not *Bacillus subtilis*, *Escherichia coli*, or *Pseudomonas aeruginosa*. The observed activity may correlate with a GQD's ability to disrupt the bacterial cell envelope and is the initial step for cell envelope disruption, suggesting the importance of GQDs' chemical composition and shapes.

The same parameters (i.e., mainly the functionalization degree, size, and shape/geometries) also influence the biocompatibility of graphene and GO derivatives. In particular, the functionalization seems to be the required/necessary condition to achieve the highest biocompatibility for nanomaterials [24] because they provide oxygenated bridges, very useful in establishing covalent and/or electrostatic interactions with biopolymers and biomolecules. This strategy focuses on the fabrication of several different biocompatible coatings; surrounding the nanostructured material surfaces, edges, and sides; and increasing their biocompatibility towards cells and tissues.

There is also a first level of functionalization degree, which is absent in biopolymers and present only in as-deposited GO and GO derivatives (this latter as post-synthesis treated carbon-based graphene materials). In this case, the only presence of the oxygenated functional groups on the GO surface mainly provides a better dispersibility in the working medium [25]. A stable nano-dispersion, in liquid phase, offers two important advantages during the experiments:

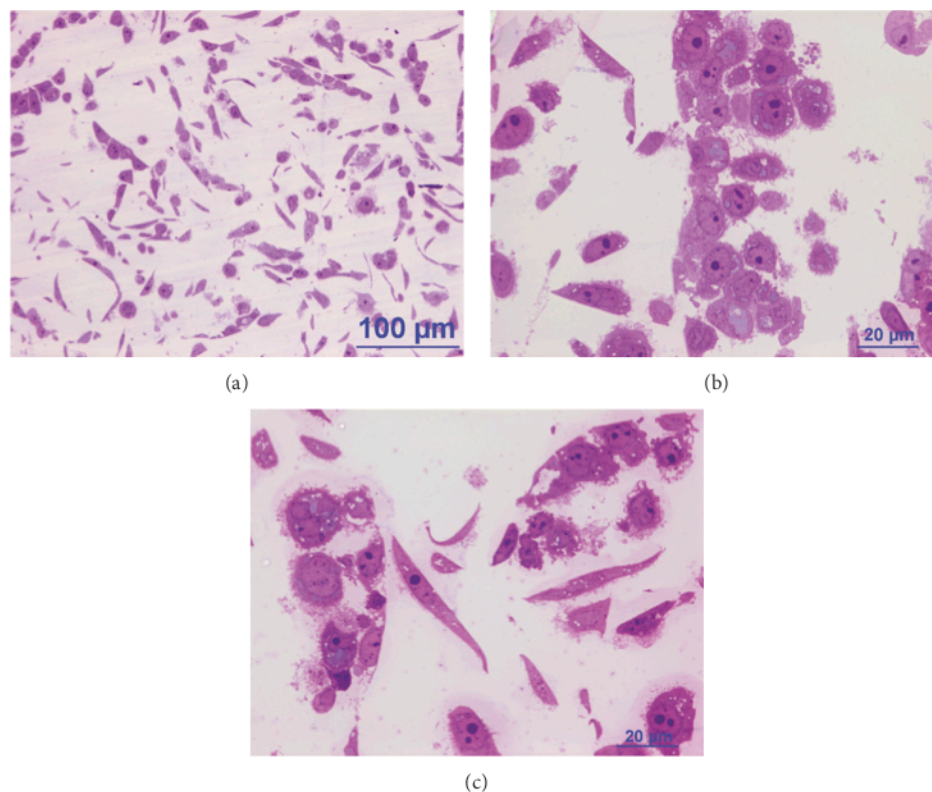


FIGURE 4: Optical microscopy of SK-N-BE(2) treated with  $2.0 \mu\text{g/mL}$  of GO +  $10 \mu\text{M}$  of S29, at (a) 24 h, magnification 20x; (b) 24 h, magnification 60x; (c) 48 h, magnification 60x (May Grünwald Giemsa staining).

- (1) Graphene nanosheets are more easily distributed in the cellular compartments, avoiding aggregation and their subsequent expulsion by gravimetric settling
- (2) Graphene nanodispersion phases improve safety bio-distribution processes [26] into the cells, minimizing the ROS generation mechanisms very well. This is especially due to the establishment of the couple electrostatic interactions between the oxygenated functional groups (mainly  $\text{C(=O)OH}$ ,  $\text{C(=O)R}$ ,  $\text{C(=O)H}$ ,  $\text{OH}$ , etc.) and the aqueous working medium. The couple interactions strongly compete with the side reactions, involved in the formation of ROS radicals, since the electrostatic interactions, established with solvent-working medium, thermodynamically stabilize the resulting nanodispersion systems [27, 28]

According to the main importance of the functionalization degree on the GO properties and performances, FTIR spectra profiles (see Figure 1), TG/DTA, and Raman results (these latter reported in Table 4) were also inserted/included in the manuscript, in order to confirm the XPS data analysis. Moreover, the acidic site evaluation (nmol/mg), reported in Table 4, mainly quantifies the  $\text{C(=O)OH}$  functional groups. These latter were also detected by using the Z-potential measurement technique that revealed negative values of  $-28.7 \text{ mV}$  for GO and  $-32.1 \text{ mV}$  for  $\text{GO}_{(\text{LiCl})}$  (typically achieved when in the presence of the negatively charged carboxylate anions  $\text{C(=O)O}^-$ ).

Furthermore, TEM analysis (Figure 2) confirmed the presence of more defective structures in the case of the most oxidized sample, as GO. This latter appeared as a multilayer of graphene oxide (A), while  $\text{GO}_{(\text{LiCl})}$  showed a more regular shape/geometry of the resulting nanosheets, similar to that exhibited by the pristine single layer of graphene.

**3.2. Cytotoxicity: Evaluation of Cellular Proliferation.** Cell growth curves in Figure 3 show that there is a decrease in growth rate during exposure to  $2 \mu\text{g/mL}$  GO for 24 and 48 hours compared to controls. Cell mortality rate, detected by trypan blue dye, was between 4% and 6%. We had previously established that higher doses of GO (10 and  $20 \mu\text{g}$ ) were cytotoxic for SK-N-BE(2) cells but  $2 \mu\text{g/mL}$  GO did not induce cell damage, so this concentration of GO provided the right concentration for carrying the minimum toxic concentration of S29. Treatment of cells with  $10 \mu\text{M}$  S29 and  $2 \mu\text{g/mL}$  GO started to inhibit cell growth after 24 h of exposure and became significantly different from GO alone or S29 alone at 48 h and 72 h.  $10 \text{ nM}$  and  $100 \text{ nM}$  S29 with, respectively,  $5.74 \text{ ng/mL}$  and  $57.4 \text{ ng/mL}$  GO failed to induce significant cell growth inhibition after 24 h from addition to cultures, meaning that S29 has a very low cytotoxicity in SK-N-BE(2) cells. In particular, these optimized concentration values for both the nanocarrier and S29 therapeutic agent were the lowest results compared to those declared in literature [31]. Moreover, the same authors [31] do not load drugs but only evaluate the cytotoxicity of the GO-Ag nanocarrier, at the

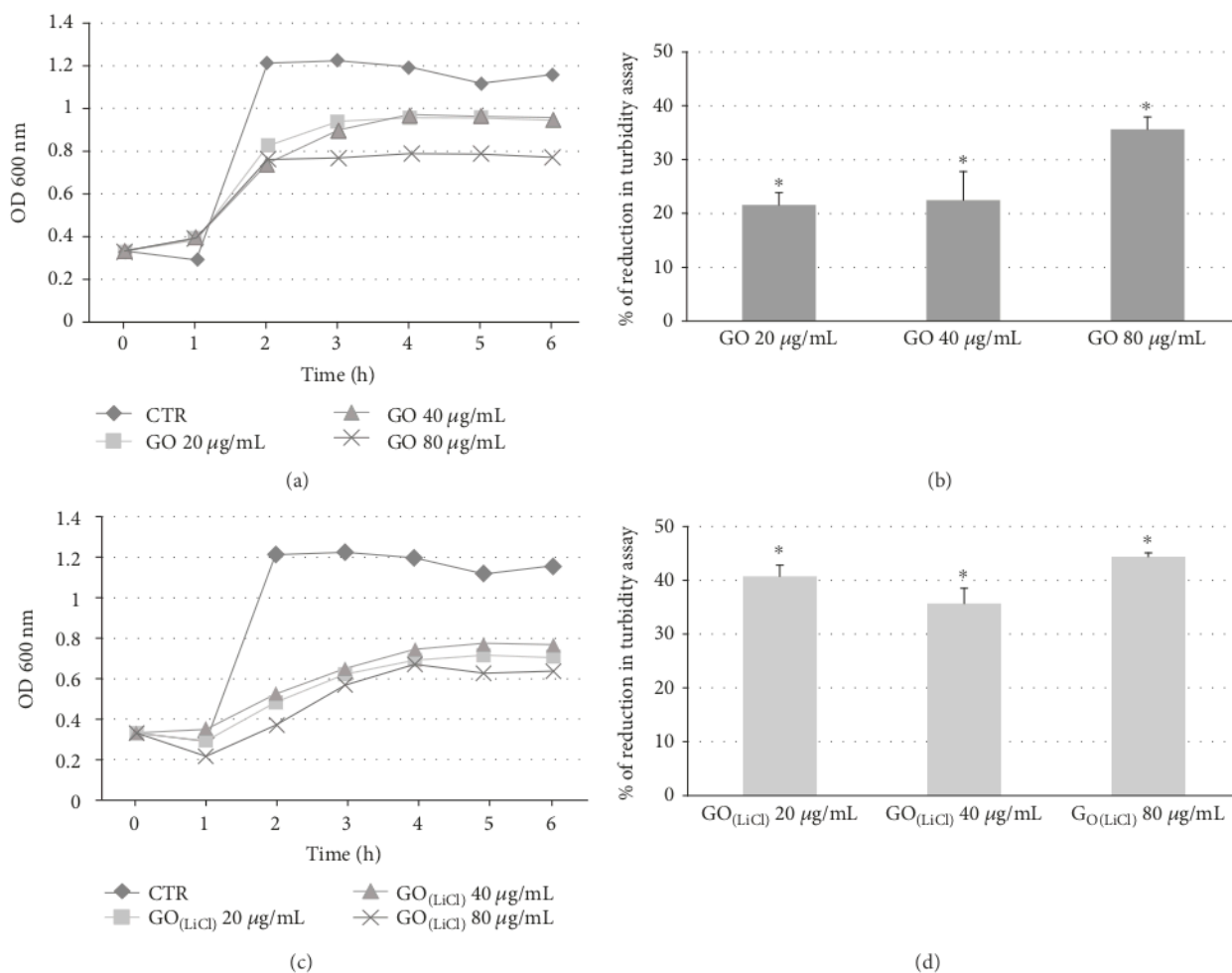


FIGURE 5: OD (a, c) and turbidimetric assay (b, d) of the two different GO incubated with *E. coli* at different times and concentrations.

highest concentration values in cells. Cell counts with trypan blue exclusion test demonstrated a low rate of cell necrosis even after 72 h of culture with  $10 \mu\text{M}$  S29 and  $2 \mu\text{g}$  GO, meaning that the inhibition of cell growth may be probably due to longer cell cycles or induction of apoptosis (the trypan blue data were also confirmed by the LDH assay, carried out as previously reported in reference [12]). Observation of cell cultures (40x inverted microscopy) showed that cells did not change their morphology nor detach from plastic after treatments (Figure 4).

**3.3. Antibacterial Activity.** Figures 5 and 6 showed the growth curves of GO-treated and  $\text{GO}_{(\text{LiCl})}$ -treated *E. coli* and *S. aureus* bacterial cells, respectively. Both treated bacteria grew following the trend exhibited by the corresponding control samples. When the bacteria cells were incubated with different graphene oxide concentrations, a significant reduction of the OD values was recorded (if compared to the controls, reported in Figures 5(b) and 5(d) and Figures 6(b) and 6(d)). Results indicated that the bacteria cells exhibited a reduction of their growth, proportionally to the GO concentrations. This agreed with the literature [32] where the lower sizes of GO sheets adhered on the bacteria surfaces, without

isolating them from the nutrient environment. Increasing the GO concentrations, bacteria cell results were more covered, effectively isolated from the environment and condemned to die. The bacteria cells, treated with  $\text{GO}_{(\text{LiCl})}$ , exhibited a significant reduction to their growth, already at the lowest concentrations and independent of the  $\text{GO}_{(\text{LiCl})}$  concentrations. This could be explained considering the largest area of  $\text{GO}_{(\text{LiCl})}$  (Table 1), which was suitable to cover the external surface of the bacteria cells, provoking the indirect toxicity by biologically isolating them from the growth medium. Consequently, the bacterial cells could neither proliferate nor consume the nutrients and died. The death rate of the bacteria was determined by the colony-counting method. Figure 7 showed that both GO and  $\text{GO}_{(\text{LiCl})}$  reduced the growth of the *S. aureus* and *E. coli* (over 90%), when bacteria were exposed to different concentrations of graphene oxides, at  $T_1 = 4$  h. The *E. coli* control samples at  $T_0$  were  $1.9 \times 10^8$  UFC/mL, and after 4 h, the control increased up to  $1.25 \times 10^9$  CFU/mL. After 4 h, the *E. coli* cells grew up to  $1.64 \times 10^8$  (at  $20 \mu\text{g/mL}$  of GO),  $1.13 \times 10^8$  (at  $80 \mu\text{g/mL}$  of GO),  $1.22 \times 10^8$  (at  $20 \mu\text{g/mL}$  of  $\text{GO}_{(\text{LiCl})}$ ), and  $1.01 \times 10^8$  CFU/mL (at  $80 \mu\text{g/mL}$  of  $\text{GO}_{(\text{LiCl})}$ ). At  $T_0$ , the control sample for *S. aureus* was  $3.8 \times 10^7$  CFU/mL, and after 4 h,

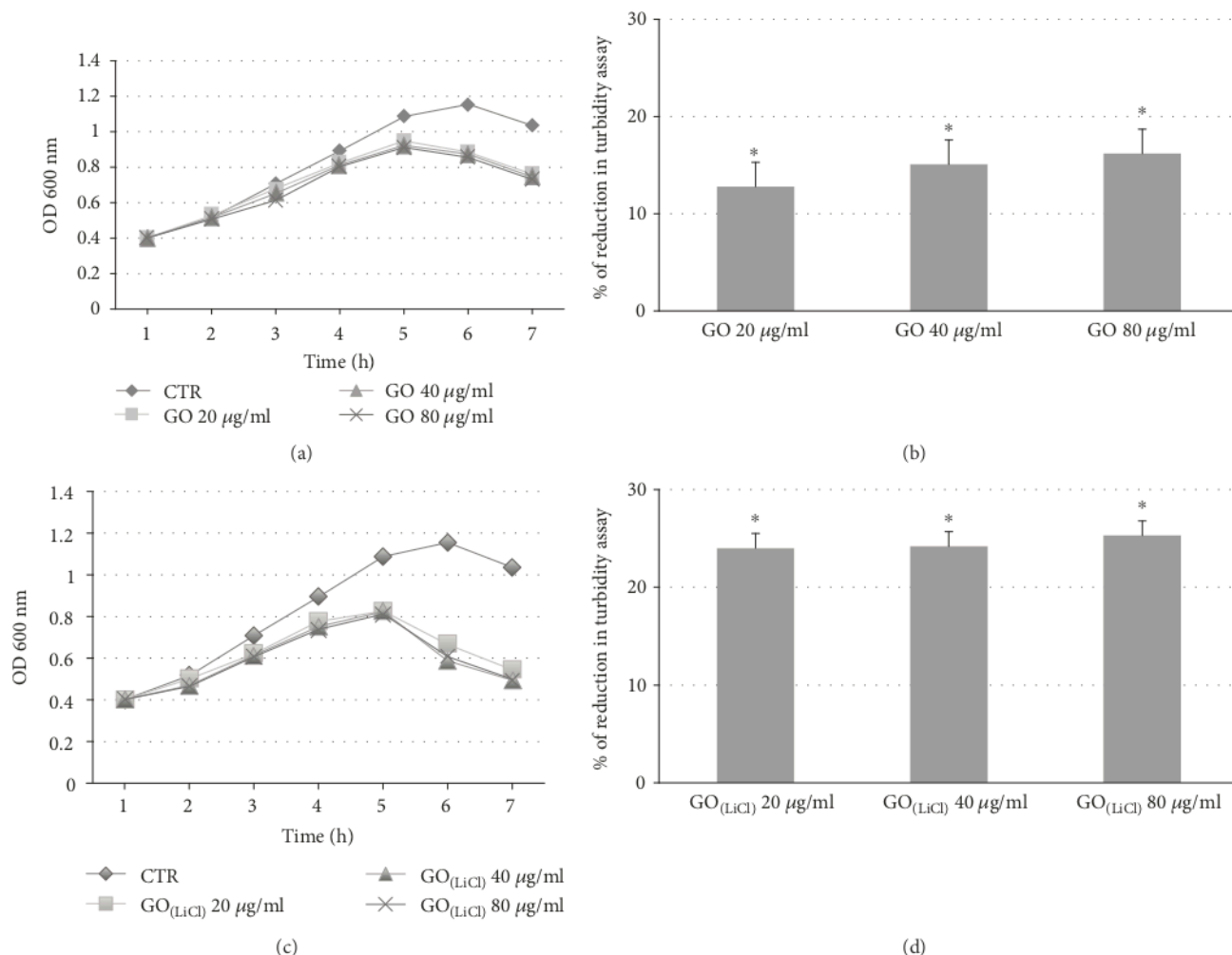


FIGURE 6: OD (a, c) and turbidimetric assay (b, d) of the two different GOs incubated with *S. aureus* at different times and concentrations.

the control increased up to  $1.53 \times 10^8$  UFC/mL. *S. aureus*-treated cells, at  $T_1$ , grew up to  $7.27 \times 10^6$  (at  $80 \mu\text{g/mL}$  of GO),  $1.85 \times 10^7$  (at  $20 \mu\text{g/mL}$  of GO),  $2.32 \times 10^6$  (at  $80 \mu\text{g/mL}$  of  $\text{GO}_{(\text{LiCl})}$ ), and  $9.64 \times 10^6$  CFU/mL (at  $20 \mu\text{g/mL}$  of  $\text{GO}_{(\text{LiCl})}$ ). After 24 hours ( $T_2$ ), the experimental results showed that *E. coli* control increased up to  $1.35 \times 10^8$ . The cells grew up to  $9.52 \times 10^7$  (at  $80 \mu\text{g/mL}$  of GO), reaching the value of  $1.03 \times 10^8$  (at  $20 \mu\text{g/mL}$  of GO). Incubating *E. coli* with  $\text{GO}_{(\text{LiCl})}$ , cells grew up to  $7.14 \times 10^7$  (at  $80 \mu\text{g/mL}$   $\text{GO}_{(\text{LiCl})}$ ) and reached the values of  $9.30 \times 10^7$  CFU/mL (at  $20 \mu\text{g/mL}$   $\text{GO}_{(\text{LiCl})}$ ). In the case of *S. aureus*, the control increased up to  $1.41 \times 10^5$ . The treated *S. aureus* cells grew up to  $5.38 \times 10^4$  (at  $80 \mu\text{g/mL}$  of GO),  $5.72 \times 10^4$  (at  $20 \mu\text{g/mL}$  of GO),  $3.91 \times 10^4$  (at  $80 \mu\text{g/mL}$  of  $\text{GO}_{(\text{LiCl})}$ ), and  $5.64 \times 10^4$  CFU/mL (at  $20 \mu\text{g/mL}$  of  $\text{GO}_{(\text{LiCl})}$ ). The antibacterial effects on *E. coli* were highly evident at 4 hours, with a growth reduction of 90%, for both GO samples. At 24 hours, the antibacterial activity for *E. coli* in the presence of  $\text{GO}_{(\text{LiCl})}$  (Figure 8(a)) was superior to GO and also higher if compared with *S. aureus*. A reasonable explanation could be the presence of a much thinner peptidoglycan wall, in the *E. coli* structure [31]. For *S. aureus*, the antibacterial effects were observed at  $T_1 = 4$  hours, with a bacterial growth reduction

of 95%. At 24 hours, the bacterial cell growth reduction was about 60%, working with GO, and slightly higher (75%) in the presence of  $\text{GO}_{(\text{LiCl})}$  (Figure 8(b)). After 24 h of GO incubation, *S. aureus* showed a typical bacteriostatic activity, if compared with *E. coli*. The experimental results could be explained as a combination of different parameters [33], such as the bacteria morphology/structure [34, 35] and the conductivity behavior of the GO [36] (more insulating nanomaterial) and  $\text{GO}_{(\text{LiCl})}$  [37] (a semiconductive material), these latter being responsible for the electron-transfer interaction between the graphene derivatives and the bacteria cells.

To examine whether direct interaction of bacteria and graphene oxide nanosheets could exhibit a key role in the antibacterial activity of graphene oxide ( $80 \mu\text{g/mL}$ ), the morphological changes of both bacterial cells, *S. aureus* and *E. coli*, were studied using SEM and TEM analysis, respectively. *E. coli* control exhibited normal rod-like morphology with regular and intact membranes (Figure 9(a)). *E. coli* incubated with GO exhibited severe membrane deformations, and cell shape was seriously distorted with evident shrinkage effects (Figure 9(b)). These effects could be probably due to a collapse of the internal structure. In some cases, the outer



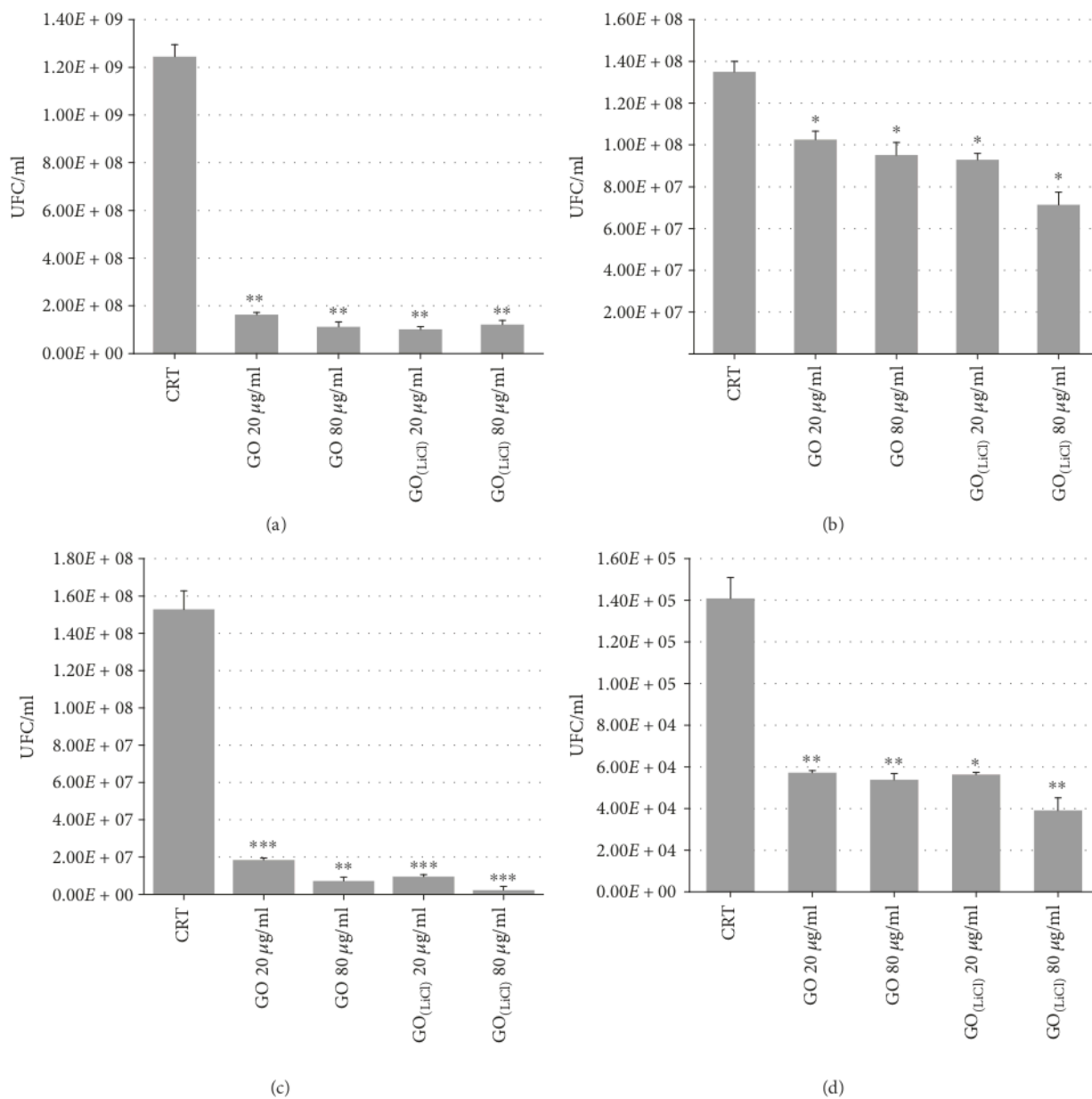


FIGURE 7: *E. coli* cell viability at 4 h (a) and 24 h (b). *S. aureus* cell viability at 4 h (c) and 24 h (d).

membranes merged with those belonging to the other bacteria cells (Figure 9(c)). *E. coli*, when incubated with GO<sub>(LiCl)</sub>, also exhibited similar behavior (Figures 9(d) and 9(e)). The experimental results could be explained considering a typical electron-transfer process from the positively charged Gram-negative *E. coli* and the negatively charged GO and GO<sub>(LiCl)</sub> nanomaterials. The results could be interpreted considering a synergy of effects such as an electron transfer [38] (from graphene oxide toward the bacterial membranes), the oxidative stress [39], and the mechanical tearing (cutting effect depending on the shape and size of the graphene nanosheets).

In the case of *S. aureus* cells, SEM micrographs showed a regular sphere shape (Figure 9(f), inset) and a typical grape-like aggregation in the entire control sample

(Figure 9(f), the main image). When *S. aureus* cells were incubated with 80  $\mu\text{g/ml}$  of GO (24 h), they showed a lightly irregular and shrunken shape (Figure 9(g)). The graphene oxide-bacteria interactions developed strong forces of attraction among cells that induced the formation of numerous spherical bacterial aggregates around graphene oxide nanosheets (Figure 9(h)), which are completely absent in the case of *E. coli*, when incubated with the same GO. This result could be explained when considering a strong electrostatic interaction between the Gram-positive *S. aureus* (with an isoelectric point [40]  $\text{pI} = 2\text{--}3$ , versus  $\text{pI} = 4\text{--}5$  in the case of *E. coli* [41]) and the more insulating GO nanomaterial. When inoculated with GO<sub>(LiCl)</sub>, the aggregation effects were also recorded (Figures 9(i) and 9(j)) in *S. aureus* cells but much less evident than those shown in Figure 9(h), for

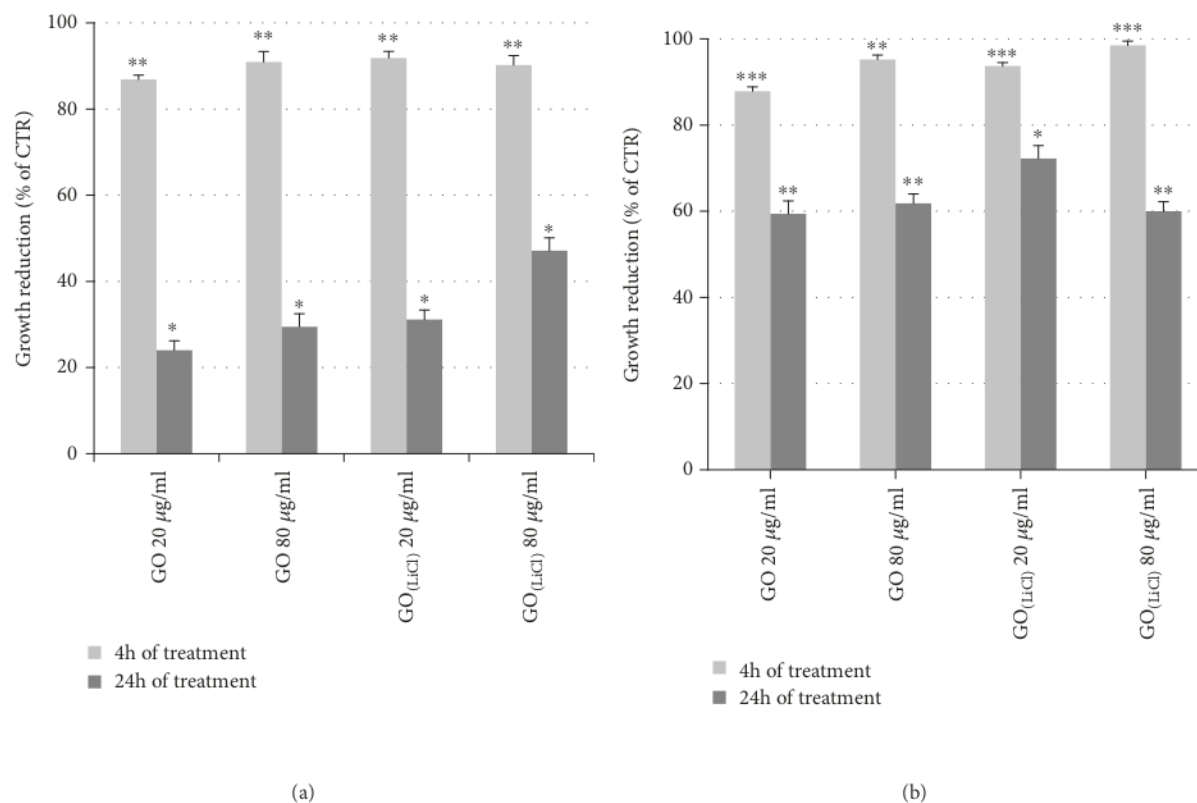


FIGURE 8: Growth reduction (%) of *E. coli* (a) and *S. aureus* (b) in the presence of GO or GO<sub>(LiCl)</sub>.

GO. The explanation could be related to the semiconductive properties of GO<sub>(LiCl)</sub> that could provoke a greater electrostatic repulsion, because it delivers more electrons toward the negatively charged bacterial walls, as in the case of the *S. aureus* pathogen (with the result of decreasing the value of the electrostatic potential membrane  $\Delta E/V$ ). TEM analysis confirmed that the direct contact of the graphene oxide nanosheets with bacterial cells induced the loss of the structural integrity, in both bacterial species and the presence of both GO samples. TEM micrographs of *E. coli* exposed to GO for 24h showed that the bacterial shrinkage, observed by SEM, was consistent with plasma membrane damage and retraction, leading to the leakage of intracellular material in the periplasmic space (Figures 10(b) and 10(c) versus Figure 10(a) that is the control). Therefore, it is reasonable to think that an electron-transfer process can be triggered, following the breaking of the membrane and consequent leakage of periplasmic materials. The effective and irreversible damages of the bacterial cells could happen after the electronic transfer processes, in combination with the oxidative stress or additional mechanical tearing processes [42]. In Figure 10(c), the “insertion/cutting mode” interaction between GO and *E. coli* clearly highlighted a nanosheet while it entered into the phospholipid bilayer, cutting the membrane (Figure 10(c)), as reported in literature [43]. Similar profiles were detected incubating *E. coli* with GO<sub>(LiCl)</sub> nanosheets (Figures 10(d) and 10(e)). *S. aureus*, incubated with 80 µg/mL of GO for 24h, showed numerous contacts with GO, with a great number of cells characterized by several

types of plasma membrane damage, having significant condensation of the intracellular material [44] (Figure 10(g) versus the unaltered control in Figure 10(f)). The initial plasma membrane damage evolved in the leakage of intracellular material and the loss of cell wall integrity (Figure 10(h)), provoking the cell death. A similar trend was observed when *S. aureus* was incubated with GO<sub>(LiCl)</sub>, where the cell damage seemed to be related to the electron diffusion mechanisms, from the graphene nanosheets toward the bacterial plasma membranes, which is especially more significant for the semimetallic graphene (GO<sub>(LiCl)</sub>) samples (Figures 10(i) and 10(j)). Concerning the antibacterial activity, it is necessary to underline that both types of graphene intrinsically exhibited this property [31] without the synergy with a traditional antibacterial agent, such as the silver element.

#### 4. Conclusions

In this study, two different metal-free graphene oxide products were synthesized, by using chemical and electrochemical procedures. The resulting graphene oxides showed different functionalization degrees (mainly as C(=O)OH groups), which determined different electronic features: a semiconductive behavior in the presence of GO<sub>(LiCl)</sub> and a more insulating profile in the case of chemically synthesized GO. These features could explain the slightly different reactivity, observed working with GO and GO<sub>(LiCl)</sub> in bacteria and in neuroblastoma cells, respectively. In particular, both graphene oxides (2 µg/mL)

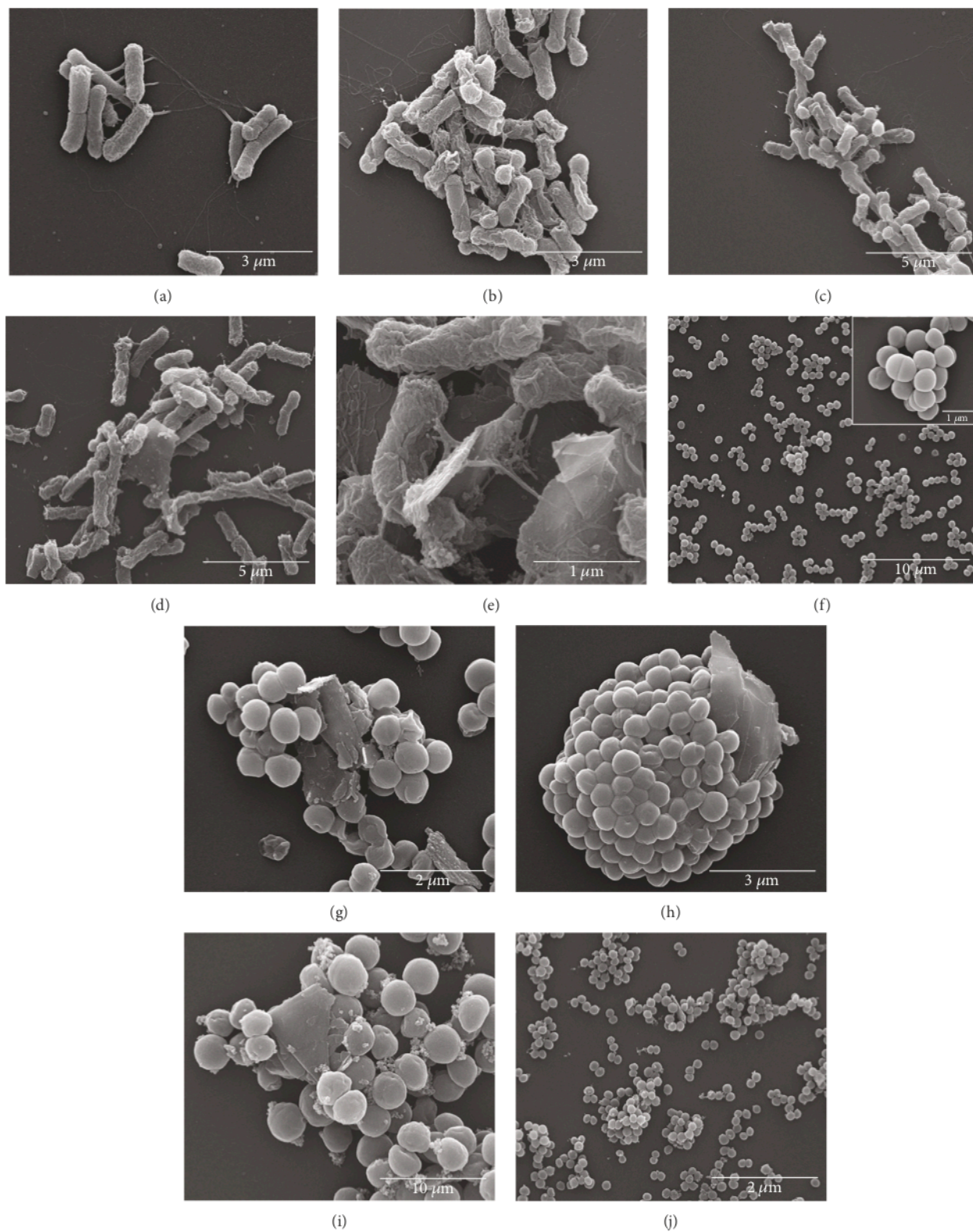
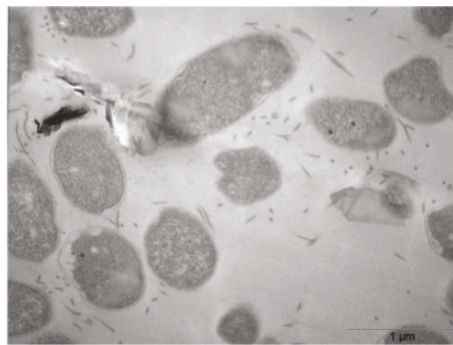


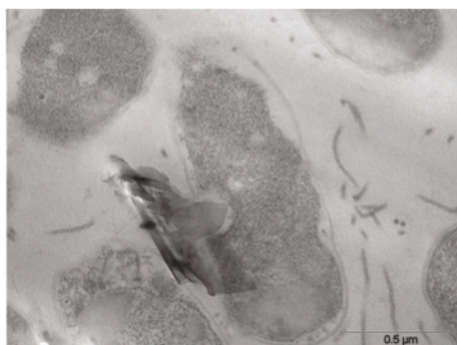
FIGURE 9: SEM micrographs of antibacterial effect of GO and  $\text{GO}_{(\text{LiCl})}$  on *E. coli* and *S. aureus*: (a) *E. coli* control; (b, c) *E. coli* incubated with  $80 \mu\text{g/mL}$  of GO at 24 h; (d, e) *E. coli* incubated with  $80 \mu\text{g/mL}$  of  $\text{GO}_{(\text{LiCl})}$  at 24 h; (f) *S. aureus* control; (g, h) *S. aureus* incubated with  $80 \mu\text{g/mL}$  of GO at 24 h; (i, j) *S. aureus* incubated with  $80 \mu\text{g/mL}$  of  $\text{GO}_{(\text{LiCl})}$  at 24 h.



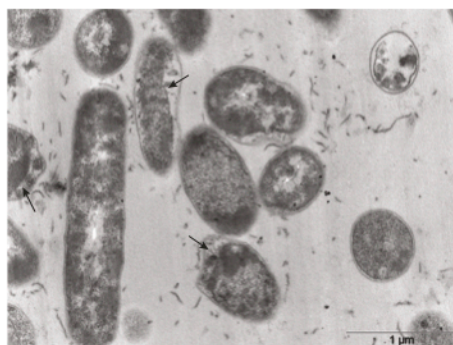
(a)



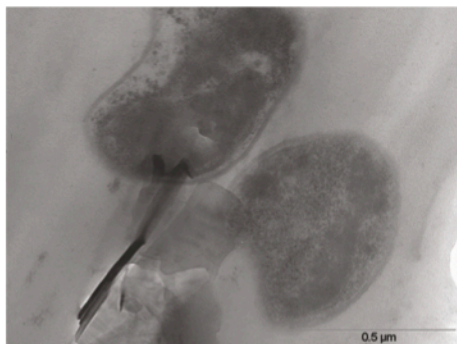
(b)



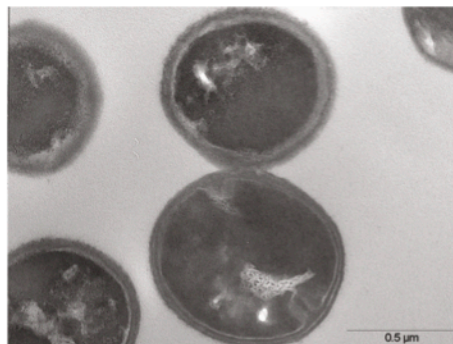
(c)



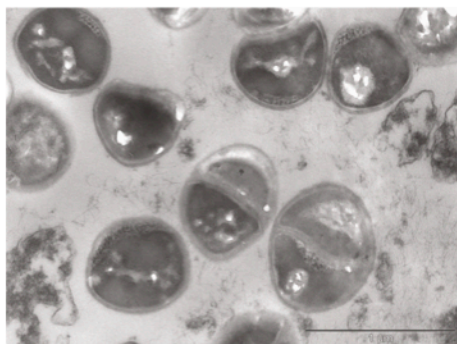
(d)



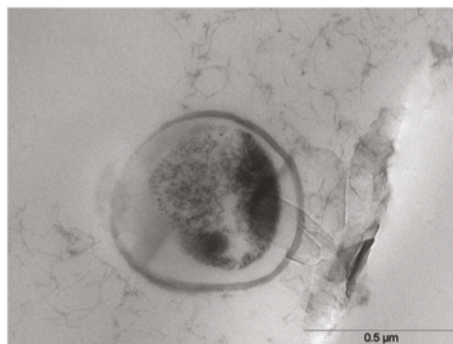
(e)



(f)



(g)



(h)

FIGURE 10: Continued.

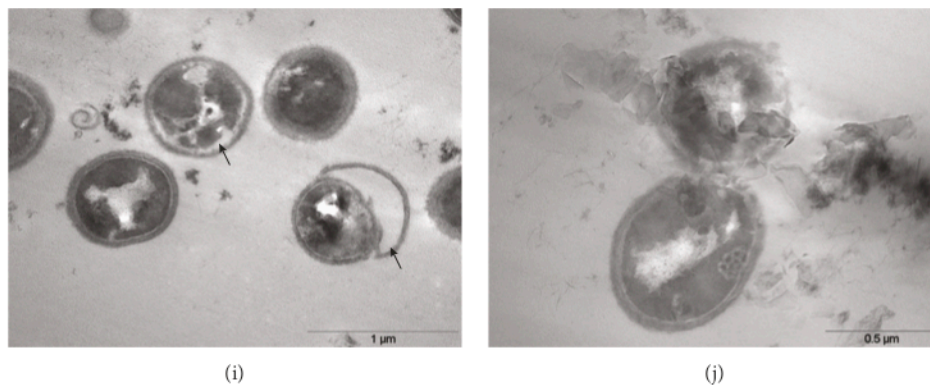


FIGURE 10: TEM micrographs of antibacterial effect of GO and  $\text{GO}_{(\text{LiCl})}$  on *E. coli* and *S. aureus*: (a) *E. coli* control; (b, c) *E. coli* incubated with  $80 \mu\text{g/mL}$  of GO at 24 h; (d, e) *E. coli* incubated with  $80 \mu\text{g/mL}$  of  $\text{GO}_{(\text{LiCl})}$  at 24 h; (f) *S. aureus* control; (g, h) *S. aureus* incubated with  $80 \mu\text{g/mL}$  of GO at 24 h; (i, j) *S. aureus* incubated with  $80 \mu\text{g/mL}$  of  $\text{GO}_{(\text{LiCl})}$  at 24 h.

showed a negligible effect of cytotoxicity, being eligible as a carrier of the S29 compound. Furthermore, both graphene oxides possess an interesting antimicrobial activity, especially  $\text{GO}_{(\text{LiCl})}$  which resulted in a bactericidal agent, and GO was a bacteriostatic ecofriendly nanomaterial. In particular, this last feature represents an important biocompatibility property that elects the GO as the new nanovector of anticancer drugs, especially in neuroblastoma cell lines.

### Data Availability

The data used to support the findings of this study are included within the article.

### Conflicts of Interest

The authors declare no competing financial interest.

### Acknowledgments

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